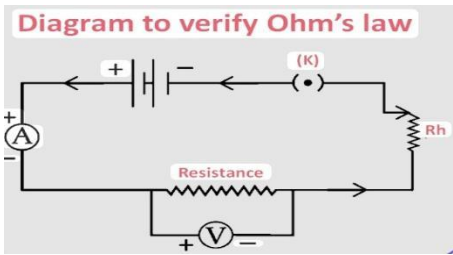


	BCM SCHOOL BASANT AVENUE DUGRI, LUDHIANA Class: X Subject: Science. Chapter- Answerkey Carbon and Its Compounds, Electricity Date: November 8,2025	
I	Multiple Choice Questions	
1	a	
2	b	
3	c	
II	Assertion–Reasoning Questions	
4	a	
5	c	
6	<p>Diagram to verify Ohm's law</p>  <p>current is directly proportional to p.d. V-I graph is straight line passing through the origin with constant slope with x axis.</p>	
7	<p>P = Ethanol (C_2H_5OH) (constituent of wine).</p> <p>Q = Ethanoic acid (CH_3COOH) (product on oxidation).</p> <p>R = Hydrogen (H_2) (gas evolved with sodium; burns with pop).</p> <p>Reactions (balanced, simplest forms): Oxidation of ethanol to ethanoic acid (with oxidising agent): $CH_3CH_2OH + [O] \rightarrow CH_3COOH + H_2O$ (in lab: acidified $K_2Cr_2O_7$ oxidises ethanol to ethanoic acid)</p>	

	<p>Reaction of ethanoic acid with sodium:</p> $2\text{CH}_3\text{COOH} + 2 \text{Na} \rightarrow 2\text{CH}_3\text{COONa} + \text{H}_2\uparrow$	
8	<p>(a)(i) Product structural formula & name: Methyl ethanoate (methyl acetate): $\text{CH}_3\text{COOCH}_3$ (structural: $\text{CH}_3\text{—C(=O)—O—CH}_3$).</p> <p>(b) Name of reaction: esterification Equation: $\text{CH}_3\text{OH} + \text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COOCH}_3 + \text{H}_2\text{O}$ (conc. H_2SO_4 as catalyst).</p> <p>(c) Why pleasant smell? Esters (like methyl ethanoate) are volatile compounds that typically have fruity/pleasant odours and interact with olfactory receptors to give a pleasant aroma.</p>	
9	<p>(a) No current will flow through the conductor as it will be open circuit.</p> <p>(b) When both key are closed</p> <p>$R' = R_1 + R_2 = 8\text{ohm}$</p> <p>$R'' = R_3 + R_4 = 8\text{ohm}$</p> <p>$R$ in parallel = 4ohm</p> <p>$I = V/R = 12/4 = 3\text{A}$</p> <p>(c) When k_1 is closed</p> <p>$R = R_1 + R_2 = 8\text{ohm}$</p> <p>$I = V/R = 12/8 = 1.5\text{A}$</p>	
10	<p>Reaction of ethanol with sodium:</p> $2 \text{C}_2\text{H}_5\text{OH} + 2 \text{Na} \rightarrow 2 \text{C}_2\text{H}_5\text{ONa} + \text{H}_2\uparrow$ <p>(product sodium ethoxide + hydrogen gas)</p> <p>(b) Why ethanol is mixed with petrol:</p>	

	<ul style="list-style-type: none"> • To raise octane rating and reduce engine knocking. • To act as an oxygenated renewable additive (reduces some harmful emissions) and partly replace fossil fuel (bioethanol). <p>(c) Two harmful effects of consuming ethanol:</p> <ul style="list-style-type: none"> • Liver damage (fatty liver → hepatitis → cirrhosis). • Central nervous system depression — impaired coordination, judgement, long-term brain damage and addiction; risk of fetal alcohol syndrome if consumed during pregnancy. 	
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